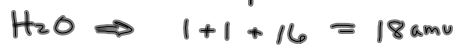
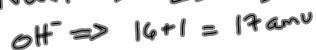
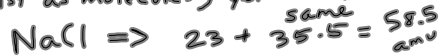


PS.198 Formula-Based Problems 8.2

Molecular mass = atomic mass of all atoms
in a molecule added up.



Formula Mass = (ionic subst. do not
exist as molecules) yet concept is the
same



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NOTE: Chemists do not figure out how
much of something by counting.
↳ find a relationship betwn.
mass & # particles.

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Relationship between masses & particles[#]

- 1 atom He = 4 amu } 2:7
 - 1 atom N = 14 amu }
 - 2 atoms He = 2 × 4 = 8 } 2:7
 - 2 atoms N = 2 × 14 = 28 }
-

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• In Lab unit for mass = grams, deliberately
found # of atoms for 1g = 1 amu

Called: Avogadro's Constant
 $6.02 \cdot 10^{23}$

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gram (g)

This # of things called mole (mol)

↳ one mole of particles
(atoms, ions, molecules) has a
mass in grams = to that
of amu

HW: 202-203
34-37

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