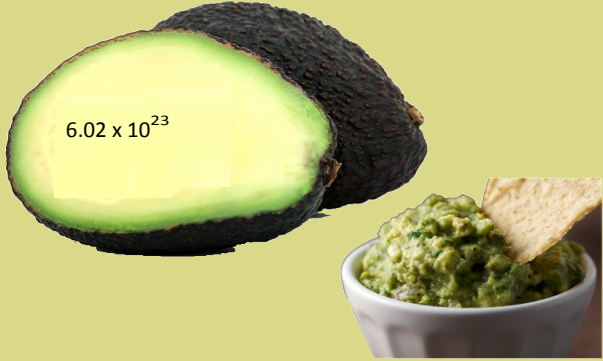


How many atoms are in 1 mole?

$6.02 \times 10^{23}$

An avocado cut in half, showing the green flesh and dark skin. Next to it is a small white bowl filled with guacamole and a single tortilla chip. The background is a light green color.

Apr 4-11:22 AM

What is the molar mass of  $\text{Na}_2\text{SiF}_6$


188g

An avocado cut in half, showing the green flesh and dark skin. Next to it is a small white bowl filled with guacamole and a single tortilla chip. The background is a light green color.

Apr 4-11:22 AM

Convert 1.098 km into m

1098 m

An avocado cut in half, showing the green flesh and dark skin. Next to it is a small white bowl filled with guacamole and a single tortilla chip. The background is a light green color.

Apr 4-11:22 AM


Convert  $491 \text{ dm}^3$  into liters

491L

An avocado cut in half, showing the green flesh and dark skin. Next to it is a small white bowl filled with guacamole and a single tortilla chip. The background is a light green color.

Apr 4-11:22 AM


Avogadro's constant represents \_\_\_\_\_ things



$6.02 \times 10^{23}$

Apr 4-11:22 AM


For ionic compounds adding the atomic masses of all the atoms in a formula unit of the compound gives the \_\_\_\_\_



formula mass

Apr 4-11:22 AM


In the copper reaction lab what was the blue precipitate in the reaction:  $\text{Cu}(\text{NO}_3)_2 + \text{NaOH} \rightarrow \text{Cu}(\text{OH})_2 + \text{NaNO}_3$



$\text{Cu}(\text{OH})_2$   
precipitates only form on the product side of equation (right of arrow) and the blue color indicates the presence of copper

Apr 4-11:22 AM

In the copper reaction lab what was the purpose of the magnesium and zinc?



To replace the Cu in the ionic solution, and release the Cu

Apr 4-11:22 AM

What happened to the Cu overnight after you added Mg?

it got absorbed back into the (aq) solution



Apr 4-11:22 AM

Find the molar mass for the following element: Ge

72.61 g



Apr 4-11:22 AM

If you have 3 moles of  $\text{H}_2\text{SO}_4$ , how many grams do you have?

$3 \times 98 = 294\text{g}$



Apr 4-11:22 AM

If you have 50 g of HCl, how many moles do you have?

1.4 mol HCl

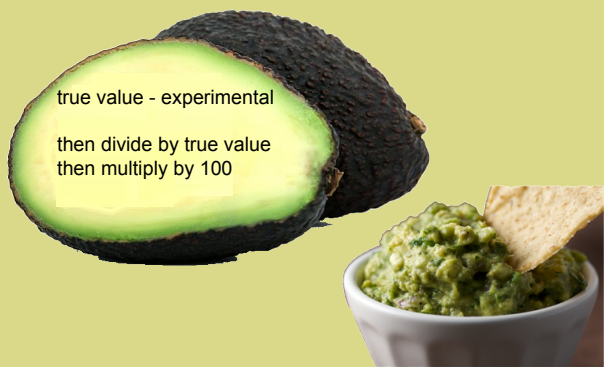


Apr 4-11:22 AM

How do you calculate percent error?

true value - experimental

then divide by true value  
then multiply by 100



Apr 4-11:22 AM